

Name: \_\_\_\_\_ SB#: \_\_\_\_\_

1. Write the oxidation number for each element on the reactant and product side of this chemical equation. (Phase labels are omitted for clarity.)



reactant side

K	
Mn	
O	
Cl	

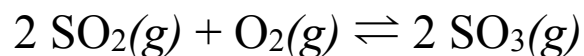
product side

K	
Mn	
O	
Cl in $\text{MnCl}_2$	
Cl in $\text{Cl}_2$	
Cl in $\text{KCl}$	

What element was oxidized? \_\_\_\_\_

What element was reduced? \_\_\_\_\_

2. Write the equilibrium constant expression for the following reaction.



\_\_\_\_\_