

Name: _____ SB#: _____

All work is to be done individually. You can use textbooks, notes, and resources on the internet.

What is the formula weight and mass % Cu for each of the following? (Show work below table.)

compound	formula weight	mass % Cu
CuCO_3		
$\text{Cu}(\text{OH})_2\text{Cu}(\text{CO}_3)$		
$\text{Cu}(\text{OH})_2\text{Cu}_2(\text{CO}_3)_2$		

When 3.8044 g $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ and 2.0503 g Na_2CO_3 were combined as described in the experimental procedure within, 1.5517 g of product was obtained. Assuming the product to be CuCO_3 , what was the percent yield of the reaction? (Show work.)